# TABLE OF ELECTROVALENCIES

| CATIONS | | ANIONS | |
|---|---|---|---|
| $Na^+$ | sodium | $Cl^-$ | chloride |
| $K^+$ | potassium | $F^-$ | fluoride |
| $Li^+$ | lithium | $Br^-$ | bromide |
| $Ag^+$ | silver | $I^-$ | iodide |
| $Cu^+$ | copper (I) | $OH^-$ | hydroxide |
| $NH_4^+$ | ammonium | $NO_3^-$ | nitrate |
| $Mg^{2+}$ | magnesium | $O^{2-}$ | oxide |
| $Ca^{2+}$ | calcium | $S^{2-}$ | sulfide |
| $Ba^{2+}$ | barium | $SO_4^{2-}$ | sulfate |
| $Zn^{2+}$ | zinc | $CO_3^{2-}$ | carbonate |
| $Cu^{2+}$ | copper (II) | | |
| $Fe^{2+}$ | iron (II) | $PO_4^{3-}$ | phosphate |
| $Pb^{2+}$ | lead | $N^{3-}$ | nitride |
| $Al^{3+}$ | aluminium | | |
| $Fe^{3+}$ | iron (III) | | |

**EXAMPLES:** Use the table of electrovalencies to determine the chemical formulae of the following compounds.

Magnesium iodide

aluminium hydroxide

sodium carbonate

iron(III) oxide

iron(III) sulfate

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**PRACTICE:** Use the table of electrovalencies to determine the chemical formulae of the following compounds.

1. sodium fluoride

2. calcium bromide

3. zinc hydroxide

4. potassium sulfate

5. barium oxide

6. silver sulfide

7. copper(II) chloride

8. magnesium nitrate

9. calcium nitride

10. magnesium phosphate